

# Chemical tests

## Mark Scheme 2

<b>Level</b>	IGCSE(9-1)
<b>Subject</b>	Chemistry
<b>Exam Board</b>	Edexcel IGCSE
<b>Module</b>	Double Award (Paper 1C)
<b>Topic</b>	Inorganic Chemistry
<b>Sub-Topic</b>	Chemical tests
<b>Booklet</b>	Mark Scheme 2

**Time Allowed:** 71 minutes

**Score:** /59

**Percentage:** /100

**Grade Boundaries:**

9	8	7	6	5	4	3	2	1
>90%	80%	70%	60%	50%	40%	30%	20%	10%

Question number			Answer	Notes	Marks
1	a	i	M1 bubbles / fizzing / effervescence OR solid/magnesium disappears/dissolves OR flask gets warm	Allow just gas (given off) Ignore wrongly named gas  Allow temperature increases but not heat produced	1
		ii	M1 magnesium chloride / MgCl <sub>2</sub>	Accept phonetic spellings Accept poorly written formulae such as MGCl <sub>2</sub> and MgCL <sub>2</sub>	1
	b	i	M1 M2 2H <sub>2</sub> + O <sub>2</sub> → 2H <sub>2</sub> O	correct formulae = 1 balancing = 1 Ignore heat anywhere Ignore state symbols	1 1
		ii	M1 condensation	Accept phonetic spellings	1
	c	i	M1 blue	Do not accept any other colours even in combination with blue, eg blue-green  Accept phonetic spellings Ignore qualifiers such as pale / dark / light Ignore mention of solution / liquid / solid	1

Question number			Answer	Notes	Marks
1	c	ii	M1 measure boiling point / melting/freezing point OR distil / boil / freeze	Ignore heat and cool	1
			M2 100 °C / 0 °C	Value must match property Accept ° or C in place of °C Do not award M2 if only value given Ignore evaporates M2 dependent on M1	1

**Total 8 marks**

Question number	Answer	Notes	Marks
2 (a) (i)	<p data-bbox="577 381 724 446"><b>H</b> × <b>H</b></p> <p data-bbox="577 487 1176 560"><b>NB</b> H does not need to be shown if touching / overlapping circles are shown</p>	<p data-bbox="1207 381 1522 479">ACCEPT any combination of dots and crosses</p> <p data-bbox="1207 519 1522 722">if overlapping / touching circles used both electrons must be within the overlapping/touching area</p>	1
(ii)	<p data-bbox="577 771 1165 868"><b>M1</b> weak forces (of attraction) between molecules / weak intermolecular forces</p> <p data-bbox="577 1015 1165 1153"><b>M2</b> (therefore) little (thermal/heat) energy required to overcome these forces / separate the molecules (into the gaseous state)</p>	<p data-bbox="1207 771 1522 974">ACCEPT particles ACCEPT bonds for forces for both <b>M1</b> and <b>M2</b> ACCEPT correctly named IMF</p> <p data-bbox="1207 1015 1522 1112">IGNORE more easily separated / easier to break</p> <p data-bbox="1207 1161 1522 1226">REJECT atoms for both <b>M1</b> and <b>M2</b></p> <p data-bbox="1207 1266 1522 1364"><b>NB</b> any mention of breaking covalent or ionic bonds scores 0</p>	2

(b)	(i)	<p><b>M1</b> <u>atoms</u> of the same element</p> <p><b>M2</b> with different masses</p>	<p><u>atoms</u> with same atomic number / <u>atoms</u> same number of protons</p> <p>different mass numbers / different numbers of neutrons</p> <p>IGNORE references to electrons unless incorrect</p>	2															
	(ii)	<table border="1"> <thead> <tr> <th></th> <th><sup>1</sup>H</th> <th><sup>2</sup>H</th> <th><sup>3</sup>H</th> </tr> </thead> <tbody> <tr> <td>protons</td> <td>1</td> <td>1</td> <td>1</td> </tr> <tr> <td>neutrons</td> <td>0</td> <td>1</td> <td>2</td> </tr> <tr> <td>electrons</td> <td>1</td> <td>1</td> <td>1</td> </tr> </tbody> </table>		<sup>1</sup> H	<sup>2</sup> H	<sup>3</sup> H	protons	1	1	1	neutrons	0	1	2	electrons	1	1	1	<p>one mark for each correct row</p>
	<sup>1</sup> H	<sup>2</sup> H	<sup>3</sup> H																
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neutrons	0	1	2																
electrons	1	1	1																

Question number	Answer	Notes	Marks
(c) (i)	exothermic		1
(ii)	$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ <b>M1</b> all formulae correct <b>M2</b> balanced	ACCEPT multiples and halves IGNORE state symbols even if incorrect	2
(iii)	<b>M1</b> (add to) <u>anhydrous/white</u> copper(II) sulfate <b>M2</b> turns blue <b>M2</b> dep on <b>M1</b> or near miss	turns copper(II) sulfate from white to blue scores 2  ACCEPT equivalent description of test with anhydrous cobalt(II) chloride (blue to pink)  IGNORE any references to testing with indicators	2

Question number	Answer	Notes	Marks
(iv)	<p><b>M1</b> <u>measure/determine</u> the boiling point</p> <p><b>M2</b> 100 °C</p> <p><b>OR</b></p> <p><b>M1</b> <u>measure/determine</u> the melting/freezing point</p> <p><b>M2</b> 0 °C</p> <p><b>OR</b></p> <p><b>M1</b> <u>measure/determine</u> the density</p> <p><b>M2</b> 1 g/cm<sup>3</sup></p>	<p>ACCEPT boil the water / heat until it boils</p> <p>it boils at 100 °C</p> <p>ALLOW "heat it and it boils at 100 °C " for 2</p> <p>ACCEPT freeze the water / cool until it freezes</p> <p>it freezes at 0 °C</p> <p>ALLOW "cool it and it freezes at 0 °C " for 2</p>	2





Question number	Answer	Notes	Marks
3 d i	white precipitate	Accept solid / ppt / ppte / suspension in place of precipitate Reject other colours Reject other observations eg fizzing Ignore cloudy/milky/grey	1
	ii silver chloride	Accept correct formula Ignore incorrect formula Award both marks if both answers in either (i) or (ii)	1
	iii (hydrochloric/sulfuric) acid / H <sup>+</sup> there OR solution acidic	Accept because there are no other ions that could form a precipitate Accept no carbonate/hydroxide (ions)	1
e	M1 wash/rinse (with water)  M2 leave it (to dry) / leave in a warm place / place in an oven / place in desiccator / heat it / dry with absorbent paper (eg kitchen/filter/blotting)	Reject methods that refer to filtrate /solution /crystallisation Ignore other named solvents  Accept leave on a window ledge Ignore evaporate it / boil it  Award 1 mark for both M1 and M2 correct but in wrong order	2
			<b>Total 10 marks</b>

Question number		Answer	Notes	Marks	
4	a	bromine: (red-)brown	Accept red	1	
		fluorine: gas AND astatine: solid	Reject orange / yellow	1	
4	b	i	Mg <sup>2+</sup> Cl <sup>-</sup>	Accept H <sup>+</sup> / H <sub>3</sub> O <sup>+</sup> Ignore OH <sup>-</sup> Award 1 for both ions correct but in wrong order	1 1
		ii	hydrogen / H <sub>2</sub> burns with a pop / squeak OR use burning/lit splint/flame to see if pop/squeak	Ignore H Must be reference to test and result Reference to splint/match with no indication of flame is not enough Reject reference to glowing splint Ignore flame extinguished 'Squeaky pop test' alone is not sufficient No ECF from wrong gas M2 DEP on M1 correct or missing	1 1
		iii	white precipitate  silver chloride to prevent other precipitates forming	Accept white solid / ppt / ppte / suspension Accept AgCl Accept to react with carbonate (ions) Accept to react with hydroxide (ions) Accept carbonates/hydroxides (also) form (white) precipitates	1 1 1

Question number			Answer	Notes	Marks
4	b	iv	hydrogen chloride/HCl does not dissociate / does not <u>form/produce</u> ions OR hydrochloric acid is not formed (in methylbenzene) /HCl is not an acid (in methylbenzene) / no H <sup>+</sup> ions (present/formed) OR only dissociates/forms H <sup>+</sup> ions in water	Accept hydrogen chloride covalent bonds do not break Ignore there is no water Ignore there are no ions (unspecified) Ignore all references to not reacting with methylbenzene	1

(Total for Question 4 = 10 marks)

Question number	Answer	Notes	Marks
5 (a) (i)	electrolysis		1
(ii)	carbon / graphite		1
(iii)	negative		1
(iv)	cryolite solvent (for alumina)  OR  to lower operating temperature / to lower melting point of mixture / electrolyte	Accept $\text{Na}_3\text{AlF}_6$ Reject to lower melting or boiling point of alumina / aluminium oxide / aluminium Ignore refs to boiling point of mixture / electrolyte Accept to reduce (heat) energy requirement Accept to increase conductivity of electrolyte <del>Ignore references to boiling point</del> Reject acts as catalyst M2 indep of M1	1 1

(b)	(i)	oxygen / O <sub>2</sub>	Ignore O	1
	(ii)	decreases capacity of blood to carry oxygen	Accept correct reference to haemoglobin / oxyhaemoglobin / carboxyhaemoglobin Accept ref to CO bonding to red blood cells but not to <u>white</u> blood cells	1
	(iii)	(pass through) limewater / calcium hydroxide solution	Ignore incorrect formulae eg CaOH Accept Ca(OH) <sub>2</sub> solution / Ca(OH) <sub>2</sub> (aq) but not just Ca(OH) <sub>2</sub>	1
		turns milky / cloudy / white	Accept chalky / white ppte etc Ignore refs to later going clear M2 dep on M1	1

5	(c)	(i)	(positive) ions / cations / Al <sup>3+</sup>	Do not accept atoms / negative ions / anions as alternative	1
			(delocalised) electrons		1
	(ii)		layers of ions/particles	Accept planes / sheets / rows Do not penalise atoms instead of ions here Reject molecules / protons / electrons	1
			slide over each other	Accept explanation in terms of non-directional bonding Do not award mark if wrong particles named, eg protons / electrons	1
	(iii)		delocalised / sea of electrons	Accept free	1
			move (through structure) / mobile	"ions free to move" scores 0	1
(iv)		low density / high strength to weight ratio	Ignore light Accept lightweight / not dense	1	

**Total 16 marks**