

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the October/November 2010 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/22

Paper 2 (AS Structured Questions), maximum raw mark 60

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Mark schemes must be read in conjunction with the question papers and the report on the examination.

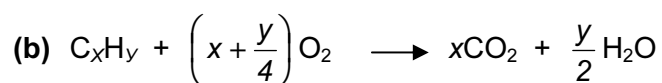
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1 (a) the actual number of atoms of each element present (1)

in one molecule of a compound (1) [2]



xCO_2 (1)

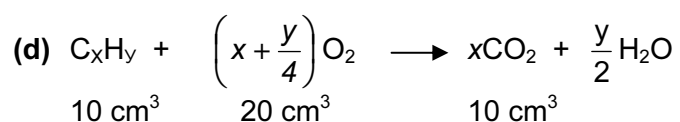
$\frac{y}{2}H_2O$ (1) [2]

(c) (i) oxygen/ O_2 (1)

(ii) carbon dioxide/ CO_2 (1)

(iii) 10 cm^3 (1)

(iv) 20 cm^3 (1) [4]



1 mol of C_xH_y gives 1 mol of CO_2

whence $x = 1$ (1)

1 mol of C_xH_y reacts with 2 mol of O_2

whence $\left(x + \frac{y}{4}\right) = 2$

and $y = 4$ (1)

molecular formula is CH_4 (1) [3]

[Total: 11]

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2 (a) $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ (1) [1]

(b) temperature between 300 and 550°C (1)

correct explanation of effect of temperature on rate of formation of NH_3 **or** on position of equilibrium (1)

catalyst of iron **or** iron oxide (1)

to speed up reaction **or** to reduce E_a (1) [4]

(c) manufacture of HNO_3

or explosives

or nylon

or as a cleaning agent

or as a refrigerant (1) [1]

(d) fertiliser in rivers causes excessive growth of aquatic plants/algae (1)

when plants and algae die O_2 is used up/fish or aquatic life die (1) [2]

(e) (i) CO by incomplete combustion of the hydrocarbon fuel (1)

NO by reaction between N_2 and O_2 in the engine (1)

(ii) CO toxic/effect on haemoglobin (1)

NO toxic/formation of acid rain (1) [4]

(f) (i) platinum/Pt – allow palladium/Pd **or** rhodium/Rh (1)

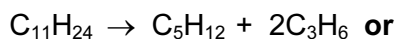
(ii) $2\text{CO} + 2\text{NO} \rightarrow 2\text{CO}_2 + \text{N}_2$ (1) [2]

[Total: 14]

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- 3 (a) (i) a compound which contains **only** carbon and hydrogen (1)
- (ii) separation of compounds by their boiling points (1) [2]

- (b) (i) high temperature **and** high pressure (1)
- high temperature **and** catalyst (1)



- (c) (i)

$CH_3CH_2CH_2CH_2CH_3$	$CH_3CH_2CH(CH_3)CH_3$	$\begin{array}{c} CH_3 \\ \\ CH_3CCH_3 \\ \\ CH_3 \end{array}$
isomer B	isomer C	isomer D
(1)	(1)	(1)

- (ii) the straight chain isomer (isomer **B** above) (1)

it has the greatest van der Waals' forces (1)

because unbranched molecules have greater area of contact/
can pack more closely together (1)

[6]

- (d) enthalpy change when 1 mol of a substance (1)

is burnt in an excess of oxygen/air under standard conditions
or is completely combusted under standard conditions (1)

[2]

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(e) (i) heat released = $m c \delta T = 200 \times 4.18 \times 27.5$ (1)

$$= 22990 \text{ J} = 23.0 \text{ kJ} \text{ (1)}$$

(ii) 23.0 kJ produced from 0.47 g of E

$$2059 \text{ kJ produced from } \frac{0.47 \times 2059}{23.0} \text{ g of E (1)}$$

$$= 42.08 \text{ g of E (1)}$$

allow ecf in (i) or (ii) on candidate's expressions [4]

(f) $C_3H_6 = 42$

E is C_3H_6

for ecf, E must be unsaturated and be no larger than C_5 (1) [1]

[Total: 18]

4	(a) reaction 1	reagent	NaOH/KOH (1)	[6]
		solvent	H_2O /water/aqueous (1)	
	reaction 2	reagent	NH_3 /ammonia (1)	
		solvent	ethanol/ C_2H_5OH /alcohol (1)	
	reaction 3	reagent	NaOH/KOH (1)	
		solvent	ethanol/ C_2H_5OH /alcohol (1)	

(b) with $CH_3CH_2CH_2CH_2I$ rate would be faster (1)

C-I bond is weaker than C-Br bond (1)

C-I bond energy is 240 kJ mol^{-1} , C-Br bond energy is 280 kJ mol^{-1}
data **must** be quoted for this mark (1) [3]

(c)	non-toxic	non-flammable	[2]
	volatile/low bp	unreactive (any 2)	

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(d) (i) when a covalent bond breaks the two electrons in the bond are shared between the two atoms (1)

(ii) $CCl_2F_2 \rightarrow CClF_2 + Cl$ (as minimum)

allow $CCl_2F + F$ (1)

[2]

(e) they are flammable (1)

[1]

[Total: 14]

5 (a) NaBr/sodium bromide

[1]

(b) Br_2 /bromine or SO_2 /sulfur dioxide

[1]

(c) concentrated sulfuric acid is an oxidising agent
or
phosphoric(V) acid is **not** an oxidising agent

[1]

[Total: 3]