

Unit 2 - Mark scheme

Question number	Answer	Mark
1	D pentane	1

Question number	Answer	Mark
2	B Y X Z W	1

Question number	Answer	Mark
3	A all Group 1 hydroxides are soluble in water	1

Question number	Answer	Mark
4	C 2-chloro-2-methylpropane	1

Question number	Answer	Mark
5	D E-1-bromo-2-methylbut-1-ene	1

Question number	Answer	Mark
6	D nucleophilic substitution	1

Question number	Answer	Mark
7	$ \begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3 - \text{C} - \text{CH}_2 - \text{O} - \text{H} \\ \\ \text{CH}_3 \end{array} $ <p>D</p>	1

Question number	Answer	Mark
8	C The solubility of hydroxides increases	1

Question number	Answer	Mark
9	B NH_4Cl	1

Question number	Answer	Mark
10	D dilute nitric acid followed by silver nitrate solution gives a white precipitate	1

Question number	Answer	Mark
11	B lithium sulfate	1

Question number	Answer	Mark
12	C emission of visible light energy as electrons return to lower energy levels	1

Question number	Answer	Mark
13	B -1 and +5	1
Question number	Answer	Mark
14	C bromine, hydrogen bromide and sulfur dioxide only	1
Question number	Answer	Mark
15(a)	A 2.5°C	1
Question number	Answer	Mark
15(b)	C redox	1
Question number	Answer	Mark
16	A an increase in the volume of acid solution	1
Question number	Answer	Mark
17(a)	A peak position shifted left, peak height higher	1
Question number	Answer	Mark
17(b)	B area A decreases, area B increases	1
Question number	Answer	Mark
18	C the forward and reverse reactions have both stopped	1

Question number	Answer	Additional guidance	Mark
19(a)	<ul style="list-style-type: none"> KBr/potassium bromide and (50%) sulfuric acid (heat under) reflux 	Both needed for M1 Ignore acid concentration Allow HBr (dry) PBr ₃ /Phosphorus(III) bromide PBr ₅ /Phosphorus(V) bromide Do not allow just heat M2 conditional on correct or near correct M1	2

Question number	Answer	Additional guidance	Mark
19(b)	<ul style="list-style-type: none"> C-Br dipole reversed OH⁻ to C arrow reversed lone pair missing (from OH⁻) 	Allow in any order	3

Question number	Answer	Additional guidance	Mark
19(c)	<ul style="list-style-type: none"> KOH/potassium hydroxide ethanol(ic)/alcohol(ic) and heat (under reflux) 	Allow NaOH/sodium hydroxide Ignore OH ⁻ / alkali M2 dependent on M1	2

Question number	Answer	Additional guidance	Mark
20(a)	<ul style="list-style-type: none"> calculation of number of moles evaluation to 2/3 SF 	Example of calculation: $0.0500 \text{ cm}^3 (\times 1000 \div 1000) = 0.0500 \text{ (mol)}$ $(0.0500 \times 90.0) = 4.50 \text{ (g)}$	2

Question number	Answer	Additional guidance	Mark
20(b)	<p>An answer that make reference to the following points:</p> <ul style="list-style-type: none"> • moles of ethanedioic acid (1) • moles of potassium hydroxide and mass of potassium hydroxide. (1) 	<p>Example of calculation:</p> <p>Moles acid = $400 \times 0.0500 \div 1000 = 2.00 \times 10^{-2}$</p> <p>Moles KOH = $2.00 \times 10^{-2} \times 2 = 4.00 \times 10^{-2}$ mol</p> <p>$4.00 \times 10^{-2} \times 56.1 = 2.24(4)$ g</p> <p>Correct answer with no working scores 2 Ignore SF except 1 SF</p>	2

Question number	Answer	Additional guidance	Mark																				
20(c)	<p>This question assesses a student's ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning.</p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table border="1" data-bbox="485 1043 769 1865"> <thead> <tr> <th>Number of indicative marking points seen in answer</th> <th>Number of marks awarded for indicative marking points</th> </tr> </thead> <tbody> <tr> <td>6</td> <td>4</td> </tr> <tr> <td>5-4</td> <td>3</td> </tr> <tr> <td>3-2</td> <td>2</td> </tr> <tr> <td>1</td> <td>1</td> </tr> <tr> <td>0</td> <td>0</td> </tr> </tbody> </table> <p>The following table shows how the marks should be awarded for structure and lines of reasoning.</p> <table border="1" data-bbox="871 1043 1358 1865"> <thead> <tr> <th></th> <th>Number of marks awarded for structure and sustained lines of reasoning</th> </tr> </thead> <tbody> <tr> <td>Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.</td> <td>2</td> </tr> <tr> <td>Answer is partially structured with some linkages and lines of reasoning.</td> <td>1</td> </tr> <tr> <td>Answer has no linkages between points and is unstructured.</td> <td>0</td> </tr> </tbody> </table>	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5-4	3	3-2	2	1	1	0	0		Number of marks awarded for structure and sustained lines of reasoning	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2	Answer is partially structured with some linkages and lines of reasoning.	1	Answer has no linkages between points and is unstructured.	0	<p>Guidance on how the mark scheme should be applied.</p> <p>The mark for indicative content should be added to the mark for lines of reasoning. For example, an answer with five indicative marking points that is partially structured with some linkages and lines of reasoning, scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p> <p>In general, it would be expected that 5 or 6 indicative points would get 2 reasoning marks, and 3 or 4 indicative points would get 1 mark for reasoning, and 0, 1 or 2 indicative points would score zero marks for reasoning.</p> <p>If there is any incorrect chemistry, deduct mark(s) from the reasoning. If no reasoning mark(s) awarded, do not deduct mark(s).</p> <p>Comment: Look for the indicative marking points first, then consider the mark for the structure of the answer and sustained line of reasoning.</p>	6
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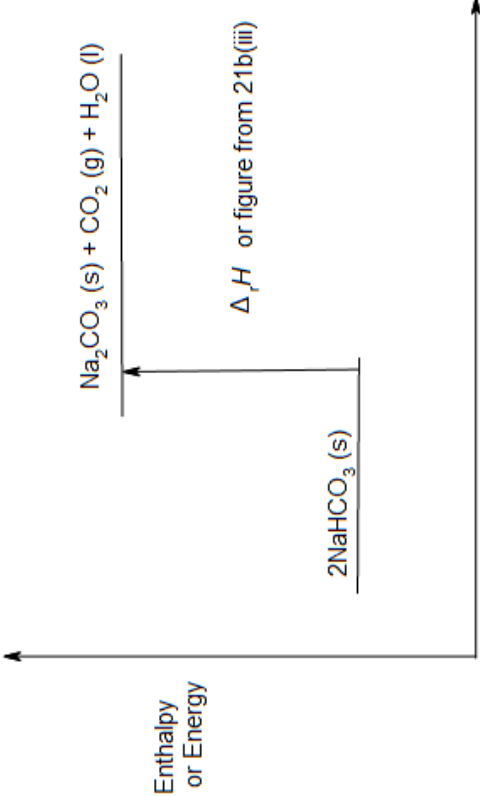
Question number	Answer	Additional guidance	Mark
20(c) Cont.	<p>Indicative points:</p> <ul style="list-style-type: none"> • rinse glassware with appropriate solutions • fill the burette with potassium hydroxide solution, ensuring there are no air bubbles • use a pipette and pipette filler to transfer 25.0 cm³ / 10 cm³ of acid to a conical flask • (add indicator to the acid in the conical flask and) carry out a range finder/rough titration • add potassium hydroxide drop by drop near the end point • repeat titrations until concordant/within ± 0.2 cm³. 	<p>Do not award just 'rinse with distilled water'. Alternative IP 2 to 5 if acid (solution) used in burette:</p> <ul style="list-style-type: none"> • fill the burette with (ethanedioic) acid solution, ensuring there are no air bubbles • use a pipette and pipette filler to transfer 25.0 cm³ of potassium hydroxide solution to a conical flask • (add indicator to the potassium hydroxide in the conical flask and) carry out a range finder/rough titration • add (ethanedioic) acid drop by drop near the end point. 	

Question number	Answer	Additional guidance	Mark
21(a)	<ul style="list-style-type: none"> • hard to measure the temperature change when you're heating something or • heat losses due to high temperatures involved or • at 300 °C/high temperatures the water will be gaseous 	<p>Allow it is difficult to measure the temperature of a solid</p>	1

Question number	Answer	Additional guidance	Mark
21(b)(i)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> • (the enthalpy change when) one mole of the substance (is formed) • from its elements in their standard states (under standard conditions). 		2

Question number	Answer	Additional guidance	Mark
21(b)(ii)	<p>A diagram that includes:</p> <ul style="list-style-type: none"> • all species correct (1) • all state symbols correct and species balanced. (1) 	$2\text{NaHCO}_3(\text{s}) \longrightarrow \text{Na}_2\text{CO}_3(\text{s}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$ <p>Do not penalise missing graphite</p>	2

Question number	Answer	Additional guidance	Mark
21(b)(iii)	<ul style="list-style-type: none"> • correct application of Hess's law (1) • correct figures used (1) • correct calculation (1) • units and sign (1) 	<p>Example of calculation:</p> $\Delta_r H = -\Delta_f H_1 + \Delta_f H_2$ <p>or</p> $\Delta_r H (\text{Na}_2\text{CO}_3) + \Delta_f H (\text{CO}_2) + \Delta_f H (\text{H}_2\text{O}) = 2\Delta_f H (\text{NaHCO}_3) + \Delta_r H$ $-1130.7 + (-285.8) + (-393.5) = 2 \times (-950.8) + \Delta_r H$ $\Delta_r H = 91.6$ $\Delta_r H = +91.6 \text{ kJ mol}^{-1}$ <p>Correct answer with no working scores (4) TE from M1 TE from incorrect M2</p>	4

Question number	Answer	Additional guidance	Mark
21(b)(iv)	<ul style="list-style-type: none"> • products energy level above reactants and arrow (1) • label on vertical arrow and vertical axis label (1) 	 <p>Enthalpy or Energy</p> <p>ΔH or figure from 21b(iii)</p> <p>(Reaction profile or progress of reaction)</p> <p>Allow reactants/products in place of chemical formulae Horizontal axis label not required Direction of arrow and endothermic/exothermic diagram must agree with sign in 21b(iii)</p> <p>Allow a correct exothermic enthalpy level diagram for an exothermic answer in 21b(iii)</p>	2

Question number	Answer	Additional guidance	Mark
22(a)(i)	<ul style="list-style-type: none"> peak in the range 3750 - 3200 cm^{-1} and O-H (stretching) bond in alcohols 	Must identify the bond and give the wavenumber range Allow peak at $\sim 3375 \text{ cm}^{-1}$	1

Question number	Answer	Additional guidance	Mark
22(a)(ii)	<ul style="list-style-type: none"> not possible - All three contain the same bonds or possible - the fingerprint regions differ/by comparing the spectra to reference spectra 	No mark for unjustified answer	1

Question number	Answer	Additional guidance	Mark
22(b)(i)	<ul style="list-style-type: none"> (all show) parent./molecular ion peak at 74 	Allow peak furthest to the right/highest m/z peak at 74 Do not award just 'peak at 74'	1


Question number	Answer	Additional guidance	Mark
22(b)(ii)	<ul style="list-style-type: none"> fragment $^+\text{CH}_3\text{CHOH} = 45$ fragment $^+\text{CH}_2\text{OH} = 31$ fragment $^+(\text{CH}_3)_2\text{COH} = 59$ 	Ignore missing charge on fragments	3

Question number	Answer	Additional guidance	Mark									
22(c)(i)	<table border="1"> <thead> <tr> <th>Organic compound used</th> <th>Name of oxidation product</th> <th>Structural formula of oxidation product</th> </tr> </thead> <tbody> <tr> <td>A</td> <td>Butanone and</td> <td>$\text{CH}_3\text{CH}_2\text{COCH}_3$</td> </tr> <tr> <td>B</td> <td>Butanal (1)</td> <td>$\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$</td> </tr> </tbody> </table>	Organic compound used	Name of oxidation product	Structural formula of oxidation product	A	Butanone and	$\text{CH}_3\text{CH}_2\text{COCH}_3$	B	Butanal (1)	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$	<p>Allow displayed or skeletal formulae</p> <p>1 mark for 2 correct names and 1 mark for each correct formula</p>	3
Organic compound used	Name of oxidation product	Structural formula of oxidation product										
A	Butanone and	$\text{CH}_3\text{CH}_2\text{COCH}_3$										
B	Butanal (1)	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$										

Question number	Answer	Additional guidance	Mark
22(c)(ii)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> • reagent: Benedict's/Fehling's (1) • (oxidation product of) compound A: no change (1) • (oxidation product of) compound B: (Benedict's/Fehling's test) red precipitate. (1) 	<p>Allow Tollens' or iodine + alkali</p> <p>(Tollens' reagent) silver mirror with (oxidation product of) B. No reaction with (oxidation product of) A</p> <p>(iodine + alkali) yellow precipitate (iodoform) with (oxidation product of) A. No reaction with oxidation product of B</p> <p>If (butanoic acid in (c)(i), allow reagent: sodium carbonate/sodium hydrogencarbonate (solution))</p> <p>Observations: (oxidation product of) compound B: bubbles/fizzes</p>	3

Question number	Answer	Additional guidance	Mark
23(a)	• balanced equation	$I_2(s) + Cl_2(g) \rightarrow 2ICl(l)$	2
	• all states correct	Accept multiples	

Question number	Answer	Additional guidance	Mark
23(b)	• correct electronegativity values and correct dipole diagram	Cl = 3.0 and I = 2.5 $\delta^+ I - Cl \delta^-$ Do not award full charges	1

Question number	Answer	Additional guidance	Mark
23(c)(i)	• 1 mark each correct formula		2
		Allow 1 mark for 2 correct non-skeletal formulae	

Question number	Answer	Additional guidance	Mark
23(c)(ii)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> • identification of correct isomer (1) • iodine is δ^+ and is attacked by the π electrons (1) • more stable secondary carbocation formed. (1) 	2-chloro-1-iodopropane	3

Question number	Answer	Additional guidance	Mark
23(d)(i)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> • carry out in fume cupboard (1) • chlorine is toxic. (1) 	<p>Allow fume hood or similar description</p> <p>Do not allow 'harmful'</p>	2

Question number	Answer	Additional guidance	Mark
23(d)(ii)	<ul style="list-style-type: none"> • I in ICl = +1 • I in ICl_3 = +3 	Both needed for the mark	1

Question number	Answer	Additional guidance	Mark
23(d)(iii)	<ul style="list-style-type: none"> • +5 and -1 to -1 (and -1) (1) • not disproportionation because the chlorine has not undergone both oxidation and reduction (1) 		2

Question number	Answer	Additional guidance	Mark
23(e)(i)	• correct method	$\text{Cl}_2 = 2 \times 35.5 = 71$ $71 \div 24000$	2
	• answer with units	$= 0.0029583 \text{ g cm}^{-3}$ $= 3 \text{ g dm}^{-3}$	

Question number	Answer	Additional guidance	Mark
23(e)(ii)	An explanation that makes reference to the following points:		3
	• chlorine (gas) is more dense than air	(1)	
	• chlorine (gas) removed (from the equilibrium)	(1)	
	• position of equilibrium moves to the LHS (more brown liquid/ICl).	(1)	

Question number	Answer	Additional guidance	Mark
23(f)	• calculation of mols of iodine and fluorine	Mols of iodine = $0.64 \div 126.9 = 5.04 \times 10^{-3}$ Mols of fluorine = $(1.31 - 0.64) \div 19 = 3.53 \times 10^{-2}$	2
	• calculation of whole number ratio and formula	Ratio 1:7 therefore formula IF_7	

