

Location Entry Codes

As part of CIE's continual commitment to maintaining best practice in assessment, CIE has begun to use different variants of some question papers for our most popular assessments with extremely large and widespread candidature. The question papers are closely related and the relationships between them have been thoroughly established using our assessment expertise. All versions of the paper give assessment of equal standard.

The content assessed by the examination papers and the type of questions are unchanged.

This change means that for this component there are now two variant Question Papers, Mark Schemes and Principal Examiner's Reports where previously there was only one. For any individual country, it is intended that only one variant is used. This document contains both variants which will give all Centres access to even more past examination material than is usually the case.

The diagram shows the relationship between the Question Papers, Mark Schemes and Principal Examiner's Reports.

| Question Paper | Mark Scheme | Principal Examiner's Report |
|-------------------------------|----------------------------|--|
| Introduction | Introduction | Introduction |
| First variant Question Paper | First variant Mark Scheme | First variant Principal Examiner's Report |
| Second variant Question Paper | Second variant Mark Scheme | Second variant Principal Examiner's Report |

Who can I contact for further information on these changes?

Please direct any questions about this to CIE's Customer Services team at: international@cie.org.uk

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

MARK SCHEME for the October/November 2008 question paper

0620 CHEMISTRY

0620/31

Paper 31 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the October/November 2008 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



| Page 2 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 31 |

- 1 red litmus paper blue [1]
OR white fumes/smoke with HCl (g) or (aq)
- chlorine [1]
- “pop” with a lighted splint or burn with a pop or goes pop and extinguishes flame [1]
NOT glowing splint
- oxygen [1]
- carbon dioxide [1]
ACCEPT correct formulae
- [Total: 5]**
- 2 (a) 3Na : 1N correct ratio [1]
correct charges [1]
8e around N [1]
- if no symbols then must have correct key
if covalent only mark 1
ignore electrons around sodium
if the response includes both a correct and an incorrect answer
do not select correct one, mark = [0]
- (b) (i) positive ions or cations [1]
NOT atoms or cores or nuclei
layers or lattice or regular pattern [1]
delocalised or free or mobile electrons or sea [1]
- OR** positive ions or cations [1]
NOT atoms or cores or nuclei
attraction between ions and electrons [1]
delocalised or free or mobile electrons or sea [1]
the attraction/electrostatic bonding must be between ions and
delocalised electrons, between cations and anions does not score
ACCEPT bond if qualified - electrostatic bond, etc.
if molecular or molecules then cannot score cation mark
- (ii) delocalised/free/mobile electrons [1]
or electrons can move [1]
- layers or ions or atoms or particles [1]
NB more flexible than 2(b)(i)
can slip or move past each other or bonding non-directional [1]

| Page 3 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 31 |

- (c) (i) tetrahedral [1]
 1Si : 4O bonded/surrounded, etc. [1]
 1O : 2 Si [1]

NOT molecules of oxygen, etc.

NOT intermolecular forces

ONLY tetrahedral can score for either of the above

Despite what the question states, **ACCEPT** a clear accurate diagram which shows the above three points.

- (ii) hard
 high mp **or** bp
 colourless (**NOT** clear) **or** shiny **or** translucent
 non/poor conductor (of electricity)
 brittle
 insoluble
 any **TWO** [2]
NOT crystalline **or** strong

[Total: 14]

- 3 (a) (i) water **or** moisture **ACCEPT** salty water [1]
 air **or** oxygen [1]

- (ii) galvanising **or** coat with zinc
 tin plate
 chromium plate
 nickel plate
 cobalt plate
 copper plate
 cover with aluminium
 anodic protection **or** sacrificial protection
 cathodic protection
 cover with plastic
 alloying (ignore any named metal)
 any **TWO** [2]
NOT just plate **or** electroplate need electroplate with suitable metal
NOT oil
ACCEPT both galvanising and sacrificial protection

- (b) (i) hydrogen **or** carbon **or** carbon monoxide **or** methane [1]
or more reactive metal **NOT** Group I

- (ii) any correct equation [2]
 only error not balanced [1]

| Page 4 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 31 |

- (c) (i) 196 [1]
- (ii) $112/196 \times 100$ [1]
 $= 57(.1)\%$ **ACCEPT** 57 to nearest whole number [1]
 mark e.c.f. to (c)(i) provided percentage not greater than 100%
ONLY ACCEPT $112/\text{answer (c)(i)} \times 100$
 otherwise [0]

- (d) (i) forms carbon dioxide/carbon monoxide (which escapes) [1]
- (ii) forms silicon(IV) oxide **or** silicon oxide **or** silica [1]
OR CaO reacts with SiO₂
 to form slag **or** calcium silicate [1]
 ignore an incorrect formula if a correct name "slag" given
NOT Si + O₂ + CaO form slag, this gains mark for slag only

[Total: 13]

- 4 (a) (i) C₆H₅COOH **or** C₆H₅CO₂H [1]
NOT C₇H₆O₂ /C₆H₆COO
- (ii) sodium hydroxide + benzoic acid = sodium benzoate + water [1]
 correct spelling needed **NOT** benzenoate
ACCEPT correct symbol equation
- (iii) sodium carbonate **or** oxide **or** hydrogencarbonate [2]
 any **TWO**
NOT Na
- (b) (i) 7.7% [1]
- (ii) for any number: equal number ratio [2]
 for example 1:1 **or** 6:6
- (iii) empirical formula is CH [1]
 molecular formula is C₆H₆ [1]
 no e.c.f., award of marks not dependent on (ii)
- (c) (i) C₆H₈O₆ [1]
- (ii) carbon – carbon double bond **or** alkene [1]
 alcohol **or** hydroxyl **or** hydroxy [1]
NOT hydroxide [1]
 hydroxide and alcohol = 0

[Total: 12]

| | | | |
|--------|-------------------------------|----------|-------|
| Page 5 | Mark Scheme | Syllabus | Paper |
| | IGCSE – October/November 2008 | 0620 | 31 |

- 5 (a) (i) $2\text{H}^+ + 2\text{e} \rightarrow \text{H}_2$ [1]
- (ii) $2\text{Cl}^- - 2\text{e} \rightarrow \text{Cl}_2$ or $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}$ [1]
- (iii) Na^+ and OH^- are left [1]
 OR Cl^- removed OH^- left
NB ions by name or formula essential
NOT any reaction of Na or Na^+
NOT Na^+ and OH^- combine
- (b) (i) sterilise/disinfect water or kill microbes/germs bacteria, etc. [1]
NOT just to make it safe to drink or purify it or clean it
 treat above as neutral they do not negate a correct response
- (ii) ammonia or methanol or hydrogen chloride or margarine [1]
NOT nylon
- (iii) fat or lipid or triester or named fat or glyceryl stearate [1]
 or vegetable oil [1]
 heat [1]
- [Total: 7]**

6 (a) (i)

| aqueous solution | tin Sn | manganese Mn | silver Ag | zinc Zn |
|-----------------------|-----------|-----------------|--------------|------------|
| tin(II) nitrate | | R | NR | R |
| manganese(II) nitrate | NR | | NR | NR |
| silver(I) nitrate | R | R | | R |
| zinc nitrate | NR | R | NR | |

[1] for each row [3]
 ignore anything written in blank space

- (ii) $\text{Sn} + 2\text{Ag}^+ \rightarrow \text{Sn}^{2+} + 2\text{Ag}$ [2]
 all species correct [1]
 accept equation with Sn^{4+}
- (iii) Mn to Mn^{2+} need both species [1]
 electron loss or oxidation number increases [1]
- (iv) covered with oxide layer [1]
 makes it unreactive or protects or aluminium oxide unreactive [1]
- (b) (i) potassium has one valency electron [1]
 or loses one electron
 calcium has two valency electrons [1]
 or loses two electrons
- (ii) potassium hydroxide \rightarrow no reaction [1]
 calcium hydroxide \rightarrow calcium oxide and water [1]
ACCEPT metal oxide

| Page 6 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 31 |

(iii) $2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$ [2]
 [1] for **formula** of either product

$2\text{Ca}(\text{NO}_3)_2 \rightarrow 2\text{CaO} + 4\text{NO}_2 + \text{O}_2$ [2]
 [1] for **formulae** of any **TWO** products

[Total: 17]

7 (a) (i) 35 cm^3 [1]
 40 cm^3 [1]

(ii) forms carbon monoxide [1]

poisonous **or** toxic **or** lethal **or** prevents blood carrying oxygen
or effect on haemoglobin [1]
NOT just harmful

(b) (i) chlorobutane **or** butyl chloride [1]
 number not required but if given must be 1, it must be in correct position

(ii) light **or** UV **or** 200°C **or** lead tetraethyl [1]

(iii) any correct equation for example 2-chlorobutane
or dichlorobutane [1]

(c) (i) correct repeat unit [1]
COND continuation [1]
 $-(\text{CH}(\text{CH}_3)-\text{CH}_2)-$

(ii) butan-1-ol **or** butan-2-ol **or** butanol [1]
 if number given then formula must correspond for second mark and number must be in correct position

structural formula of above [1]
 $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2\text{OH}$ **or** $\text{CH}_3-\text{CH}(\text{OH})-\text{CH}_2-\text{CH}_3$
NOT $\text{C}_4\text{H}_9\text{OH}$
 if first mark not awarded then either formula will gain mark [1]
ACCEPT either formula for "butanol"

(iii) $\text{CH}_3-\text{CH}(\text{Cl})-\text{CH}_3$ **or** $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{Cl}$ [1]
NOT $\text{C}_3\text{H}_7\text{Cl}$
 response must not include HCl
 if equation given look at RHS only

[Total: 12]

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

MARK SCHEME for the October/November 2008 question paper

0620 CHEMISTRY

0620/32

Paper 32 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the October/November 2008 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



| Page 2 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 32 |

- 1 ammonia [1]
 chlorine [1]
 “pop” with a lighted splint **or** burn with a pop **or** goes pop and extinguishes flame [1]
NOT glowing splint
 relights a glowing splint [1]
 turns limewater milky/cloudy/chalky/white [1]
ACCEPT correct formulae

[Total: 5]

- 2 (a) 2Na : 1S correct ratio [1]
 correct charges [1]
 8e around S [1]

if no symbols then must have correct key
 if covalent only mark 1
 ignore electrons around sodium
 if the response includes both a correct and an incorrect answer
 do not select correct one, mark = [0]

- (b) (i) positive ions **or** cations [1]
NOT atoms **or** cores **or** nuclei
 layers **or** lattice **or** regular pattern [1]
 delocalised **or** free **or** mobile electrons **or** sea [1]

OR positive ions **or** cations [1]
NOT atoms **or** cores **or** nuclei
 attraction between ions and electrons [1]
 delocalised **or** free **or** mobile electrons **or** sea [1]
 the attraction/electrostatic bonding must be between ions and
 delocalised electrons, between cations and anions does not score
ACCEPT bond if qualified e.g. electrostatic bond, etc.
 if moles or molecular cannot score cation mark

- (ii) delocalised/free/mobile electrons
or electrons can move [1]
 layers **or** ions **or** atoms **or** particles [1]
NB more flexible than 2(b)(i)
 can slip **or** move past each other **or** bonding non-directional [1]

| Page 3 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 32 |

- (c) (i) tetrahedral [1]
 1Si : 4O bonded/surrounded, etc. [1]
 1O : 2 Si [1]

NOT molecules of oxygen, etc.

NOT intermolecular forces

ONLY tetrahedral can score for either of the above

Despite what the question states, **ACCEPT** a clear accurate diagram which shows the above three points.

- (ii) hard
 high melting point **or** boiling point
 colourless (**NOT** clear) **or** shiny **or** translucent
 non/poor conductor (of electricity)
 brittle
 insoluble
 any **TWO** [2]
NOT crystalline **or** strong

[Total: 14]

- 3 (a) (i) water **or** moisture **ACCEPT** salty water [1]
 air **or** oxygen [1]

- (ii) galvanising **or** coat with zinc
 tin plate
 chromium plate
 nickel plate
 cobalt plate
 copper plate
 cover with aluminium
 anodic protection **or** sacrificial protection
 cathodic protection
 cover with plastic
 alloying (ignore any named metal)
 any **TWO** [2]
NOT just plate **or** electroplate need electroplate with suitable metal
NOT oil
ACCEPT both galvanising and sacrificial protection

- (b) (i) hydrogen **or** carbon **or** carbon monoxide **or** methane [1]
or more reactive metal **NOT** Group I

- (ii) any correct equation [2]
 only error not balanced [1]

| Page 4 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 32 |

- (c) (i) 196 [1]
- (ii) $36/196 \times 100$ [1]
 = 18(.4)% **ACCEPT** 18 to nearest whole number [1]
 mark e.c.f. to (c)(i) provided percentage not greater than 100%
ONLY ACCEPT $36/\text{answer (c)(i)} \times 100$
 otherwise [0]

- (d) (i) forms carbon dioxide/carbon monoxide (which escapes) [1]
- (ii) forms silicon(IV) oxide **or** silicon oxide **or** silica [1]
OR CaO reacts with SiO₂
 to form slag **or** calcium silicate [1]
 ignore an incorrect formula if a correct name given
NOT Si + O₂ + CaO form slag

[Total: 13]

- 4 (a) (i) C₆H₅COOH **or** C₆H₅CO₂H [1]
NOT C₇H₆O₂ /C₆H₆COO
- (ii) sodium hydroxide + benzoic acid = sodium benzoate + water [1]
 correct spelling needed **NOT** benzenoate
ACCEPT correct symbol equation
- (iii) sodium carbonate **or** oxide **or** hydrogencarbonate [2]
 any **TWO**
NOT Na
- (b) (i) 7.7% [1]
- (ii) for any number: equal number ratio [2]
 for example 1:1 or 6:6
- (iii) empirical formula is CH [1]
 molecular formula is C₆H₆ [1]
 no e.c.f., award of marks not dependent on (ii)
- (c) (i) C₆H₈O₆ [1]
- (ii) carbon – carbon double bond **or** alkene [1]
 alcohol **or** hydroxyl **or** hydroxy [1]
NOT hydroxide
 hydroxide and alcohol = 0

[Total: 12]

| | | | |
|--------|-------------------------------|----------|-------|
| Page 5 | Mark Scheme | Syllabus | Paper |
| | IGCSE – October/November 2008 | 0620 | 32 |

- 5 (a) (i) $2\text{H}^+ + 2\text{e} \rightarrow \text{H}_2$ [1]
- (ii) $2\text{Cl}^- - 2\text{e} \rightarrow \text{Cl}_2$ or $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}$ [1]
- (iii) Na^+ and OH^- are left [1]
OR Cl^- removed OH^- left
NB ions by name or formula essential
NOT any reaction of Na or Na^+
NOT Na^+ and OH^- combine
- (b) (i) sterilise/disinfect water or kill microbes/germs bacteria, etc. [1]
NOT just to make it safe to drink or purify it or clean it
 treat above as neutral they do not negate a correct response
- (ii) ammonia or methanol or hydrogen chloride or margarine [1]
NOT nylon
- (iii) ester or triester or lipid [1]
 hydrolysis or saponification [1]

[Total: 7]

6 (a) (i)

| aqueous solution | tin Sn | manganese Mn | silver Ag | zinc Zn |
|-----------------------|-----------|-----------------|--------------|------------|
| tin(II) nitrate | | R | NR | R |
| manganese(II) nitrate | NR | | NR | NR |
| silver(I) nitrate | R | R | | R |
| zinc nitrate | NR | R | NR | |

[1] for each row [3]
 ignore anything written in blank space

- (ii) $\text{Zn} + 2\text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$ [2]
 all species correct [1]
 accept correct ionic equation
 $\text{Zn} + 2\text{Ag}^+ \rightarrow \text{Zn}^{2+} + 2\text{Ag}$ [2]
- (iii) Sn^{2+} must be made clear that the oxidant is Sn^{2+} not Sn [1]
 it gains electrons or oxidation number decreases or it is reduced [1]
 reason must relate to an oxidant
NB not dependent on identifying Sn^{2+}
- (iv) covered with oxide layer [1]
 makes it unreactive or protects or aluminium oxide unreactive [1]

| Page 6 | Mark Scheme | Syllabus | Paper |
|--------|-------------------------------|----------|-------|
| | IGCSE – October/November 2008 | 0620 | 32 |

- (b) (i) potassium has one valency electron [1]
 or loses one electron
 calcium has two valency electrons
 or loses two electrons [1]
- (ii) potassium hydroxide → no reaction [1]
 calcium hydroxide → calcium oxide and water [1]
ACCEPT metal oxide
- (iii) $2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$ [2]
 [1] for **formula** of either product
- $2\text{Ca}(\text{NO}_3)_2 \rightarrow 2\text{CaO} + 4\text{NO}_2 + \text{O}_2$ [2]
 [1] for **formulae** of any **TWO** products
- [Total: 17]**
- 7 (a) (i) 20 cm^3 [1]
 80 cm^3 [1]
- (ii) forms carbon monoxide [1]
 poisonous or toxic or lethal or prevents blood carrying oxygen
 or effect on haemoglobin [1]
NOT just harmful, etc.
- (b) (i) chlorobutane or butyl chloride [1]
 number not required but if given must be 1, it must be in correct position
- (ii) light or UV or $200\text{ }^\circ\text{C}$ or lead tetraethyl [1]
- (iii) any correct equation for example 2-chlorobutane
 or dichlorobutane
 must include HCl [1]
- (c) (i) correct repeat unit [1]
COND continuation [1]
 $-(\text{CH}(\text{CH}_3)-\text{CH}_2)-$
- (ii) propan-1-ol or propan-2-ol or propanol [1]
 if number given then formula must correspond for second mark.
 number must be in correct position
 structural formula of above [1]
 $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{OH}$ or $\text{CH}_3-\text{CH}(\text{OH})-\text{CH}_3$
NOT $\text{C}_3\text{H}_7\text{OH}$
 if first mark not awarded then either formula will gain mark [1].
accept either formula for "propanol" in (i)
NB On scoris both marks entered together not as [1] and [1] separately
- (iii) $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{Cl}$ or $\text{CH}_3-\text{CH}_2-\text{CH}(\text{Cl})-\text{CH}_3$ [1]
NOT $\text{C}_4\text{H}_9\text{Cl}$
 if equation given look at RHS only
 response must not include HCl

[Total: 12]