

# Cambridge IGCSE™

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## CHEMISTRY

0620/42

Paper 4 Theory (Extended)

February/March 2021

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **12** pages. Any blank pages are indicated.



1 The table shows the numbers of protons, neutrons and electrons in particles **A** to **I**.

particle	protons	neutrons	electrons
<b>A</b>	1	0	0
<b>B</b>	6	6	6
<b>C</b>	6	8	6
<b>D</b>	10	10	10
<b>E</b>	16	16	18
<b>F</b>	17	18	17
<b>G</b>	18	22	18
<b>H</b>	19	20	19
<b>I</b>	20	20	18

Answer the following questions about particles **A** to **I**. Each letter may be used once, more than once or not at all.

(a) State which of the particles **A** to **I**:

- (i) is an anion ..... [1]
- (ii) are cations ..... and ..... [2]
- (iii) are noble gas atoms ..... and ..... [2]
- (iv) is a halogen atom ..... [1]
- (v) is a Group I atom ..... [1]
- (vi) have the same nucleon number ..... and ..... [1]
- (vii) causes acidity in aqueous solutions ..... [1]
- (viii) is used to define the relative atomic mass of elements. .... [1]

(b) Explain why **B** and **C** are isotopes of the same element.

.....

..... [2]

[Total: 12]

2 The elements shown are gases at room temperature and pressure.

hydrogen  
nitrogen  
oxygen  
chlorine

(a) State which **one** of these gases is green.

..... [1]

(b) The gases shown exist as diatomic molecules.

State the name of **another** element which has diatomic molecules and is a gas at room temperature and pressure.

..... [1]

(c) When separate samples of each of these gases are placed in a container they will diffuse.

(i) Describe why these gases diffuse.

..... [1]

(ii) State which of these four gases has the highest rate of diffusion.

Explain your answer.

gas .....

explanation .....

..... [2]

(d) Nitrogen, oxygen and other substances are found in clean, dry air.

(i) State the percentage of nitrogen in clean, dry air.

..... [1]

(ii) Other than nitrogen and oxygen, identify another element found in clean, dry air.

..... [1]

(iii) Identify a compound found in clean, dry air.

..... [1]

(iv) Nitrogen and oxygen can be separated from liquid air.

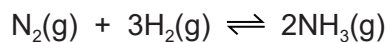
State the name of this process.

..... [2]

[Total: 10]

3 This question is about ammonia.

(a) Nitrogen reacts with hydrogen to form ammonia in an industrial process.



(i) Name this industrial process.

..... [1]

(ii) State the meaning of the symbol  $\rightleftharpoons$ .

..... [1]

(iii) State the conditions used in this industrial process. Include units.

temperature .....

pressure .....

[2]

(iv) Name the catalyst used in this industrial process.

..... [1]

(v) If the pressure is increased, the yield of ammonia increases.

Explain why, in terms of equilibrium.

.....  
 .....  
 ..... [2]

(vi) If the temperature is increased, the rate of reaction increases.

Explain why, in terms of particles.

.....  
 .....  
 .....  
 .....  
 ..... [3]

(b) Ammonia reacts with sulfuric acid to make a compound which is used as a fertiliser.

Write the chemical equation for the reaction between ammonia and sulfuric acid.

..... [2]

[Total: 12]

4 A student wanted to make some zinc chloride crystals.

The student followed the procedure shown.

**step 1** Add excess zinc powder to dilute hydrochloric acid to form aqueous zinc chloride.

**step 2** Remove unreacted zinc powder from the aqueous zinc chloride.

**step 3** Heat the solution until it is saturated.

**step 4** Allow the saturated solution to cool and remove the crystals that form.

(a) Write the equation for the reaction in **step 1**. Include state symbols.

..... [3]

(b) Explain why **excess** zinc powder is added in **step 1**.

..... [1]

(c) Suggest how unreacted zinc powder is removed in **step 2**.

..... [1]

(d) A saturated solution is formed in **step 3**.

Suggest what is meant by the term *saturated solution*.

.....  
..... [2]

(e) Explain why crystals form as the solution cools in **step 4**.

..... [1]

(f) Name **two** zinc compounds which react with dilute hydrochloric acid to form zinc chloride.

.....  
..... [2]

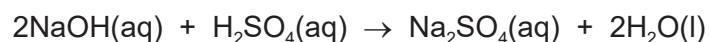
(g) If excess calcium metal is used instead of excess zinc powder in **step 1**, pure calcium chloride crystals do **not** form.

Explain why.

.....  
..... [1]

(h) Some salts can be made by titration.

In a titration experiment, 20.0 cm<sup>3</sup> of aqueous sodium hydroxide reacts exactly with 25.0 cm<sup>3</sup> of 0.100 mol/dm<sup>3</sup> dilute sulfuric acid to make sodium sulfate.



(i) Circle the name of the type of reaction that takes place.

**decomposition      neutralisation      precipitation      reduction**

[1]

(ii) Calculate the concentration of the aqueous sodium hydroxide in g/dm<sup>3</sup> using the following steps.

- Calculate the number of moles of dilute sulfuric acid used.

..... mol

- Determine the number of moles of sodium hydroxide which react with the dilute sulfuric acid.

..... mol

- Calculate the concentration of the aqueous sodium hydroxide in mol/dm<sup>3</sup>.

..... mol/dm<sup>3</sup>

- Calculate the concentration of the aqueous sodium hydroxide in g/dm<sup>3</sup>.

..... g/dm<sup>3</sup>

[5]

[Total: 17]

**QUESTION 5 STARTS ON THE NEXT PAGE.**

5 The table shows the names or structures of organic compounds **P** to **U**.

<b>P</b>	<b>Q</b>	<b>R</b>
$  \begin{array}{c}  \text{H} \quad \text{H} \quad \text{H} \\    \quad   \quad   \\  \text{H}-\text{C}-\text{C}-\text{C}-\text{H} \\    \quad   \quad   \\  \text{H} \quad \text{H} \quad \text{H}  \end{array}  $	propanoic acid	but-1-ene
<b>S</b>	<b>T</b>	<b>U</b>
propan-1-ol	methyl butanoate	$  \begin{array}{c}  \text{H} \quad \text{H} \quad \text{H} \\  \diagdown \quad   \quad   \\  \text{C}=\text{C}-\text{C}-\text{H} \\  \diagup \quad \quad   \\  \text{H} \quad \quad \text{H}  \end{array}  $

(a) Give the letters of the organic compounds, **P** to **U**, that are unsaturated hydrocarbons.

..... [2]

(b) Describe the test for an unsaturated hydrocarbon.

test .....

observations .....

[2]

(c) But-1-ene is an unbranched molecule.

(i) Name the unbranched isomer of but-1-ene.

..... [1]

(ii) Draw the structure of a branched isomer of but-1-ene. Show all of the atoms and all of the bonds.

[1]

(d) Dodecane is an alkane with 12 carbon atoms. Dodecane can be cracked.

(i) Write the formula of dodecane.

..... [1]

(ii) Give the letters of all the organic compounds, **P** to **U**, that can be formed when dodecane is cracked.

..... [2]



- (e) Name the reagent and suggest the conditions needed to convert organic compound **U** into organic compound **S**.

reagent .....

conditions .....

[3]

- (f) Organic compound **S** can be converted to organic compound **Q** by reaction with an acidified reagent.

- (i) Name the type of chemical change that happens to organic compound **S**.

..... [1]

- (ii) Name the acidified reagent added to organic compound **S**.

..... [1]

- (g) Organic compound **T** is made by reacting two compounds together.

- (i) Name the homologous series that organic compound **T** belongs to.

..... [1]

- (ii) Name the **two** compounds which react together to make organic compound **T**.

Draw the structures of each compound you have named. Show all of the atoms and all of the bonds.

name .....

structure

name .....

structure

[4]

- (iii) Deduce the molecular formula of organic compound **T**.

..... [1]

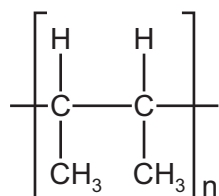
[Total: 20]

6 Polymers are large molecules built up from small molecules.

(a) State the name given to the small molecules from which polymers are made.

..... [1]

(b) The formula of a polymer is shown.



(i) Draw the structure of the small molecule from which this polymer is made. Show all of the atoms and all of the bonds.

[2]

(ii) State the type of polymerisation used to make this polymer.

..... [1]

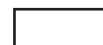
(c) Three amino acids are shown. They combine to form part of a natural polymer.



(i) Name the type of natural polymer formed when amino acids combine.

..... [1]

(ii) Complete the diagram to show part of the structure of the natural polymer that forms when these three amino acids combine. Show all of the bonds in the linkages.



[3]

(iii) Name the type of chemical reaction that takes place when this natural polymer is converted back to amino acids.

..... [1]

[Total: 9]

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The Periodic Table of Elements

		Group							
I	II	III	IV	V	VI	VII	VIII		
1	2	3	4	5	6	7	8	9	10
H hydrogen 1	He helium 4	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20		
11	12	13	14	15	16	17	18		
Na sodium 23	Mg magnesium 24	Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40		
19	20	21	22	23	24	25	26	27	28
K potassium 39	Ca calcium 40	Sc scandium 45	Ti titanium 48	V vanadium 51	Cr chromium 52	Mn manganese 55	Fe iron 56	Co cobalt 59	Ni nickel 59
37	38	39	40	41	42	43	44	45	46
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 91	Nb niobium 93	Mo molybdenum 96	Tc technetium —	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106
55	56	57–71	72	73	74	75	76	77	78
Cs caesium 133	Ba barium 137	lanthanoids	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195
87	88	89–103	104	105	106	107	108	109	110
Fr francium —	Ra radium —	actinoids	Rf rutherfordium —	Db dubnium —	Sg seaborgium —	Bh bohrium —	Hs hassium —	Mt meitnerium —	Ds darmstadtium —
81	82	83	84	85	86	87	88	89	90
Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium —	At astatine —	Rn radon —	Fr francium —	Ra radium —	Ac actinium —	Th thorium 232
91	92	93	94	95	96	97	98	99	100
Pa protactinium 231	U uranium 238	Np neptunium —	Pu plutonium —	Am americium —	Cm curium —	Bk berkelium —	Cf californium —	Es einsteinium —	Fm fermium —
101	102	103	104	105	106	107	108	109	110
Md mendelevium —	No nobelium —	Lr lawrencium —	Lu lutetium 175	Yb ytterbium 173	Tm thulium 169	Er erbium 167	Hf hafnium 168	Ta tantalum 182	W tungsten 184
111	112	113	114	115	116	117	118	119	120
Rg roentgenium —	Cn copernicium —	Nh nihonium —	Fl flerovium —	Mc moscovium —	Lv livermorium —	Ts tennessium —	Og oganeson —	Uu unbinilium —	Uub unbinilium —

Key

atomic number  
atomic symbol  
name  
relative atomic mass

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).